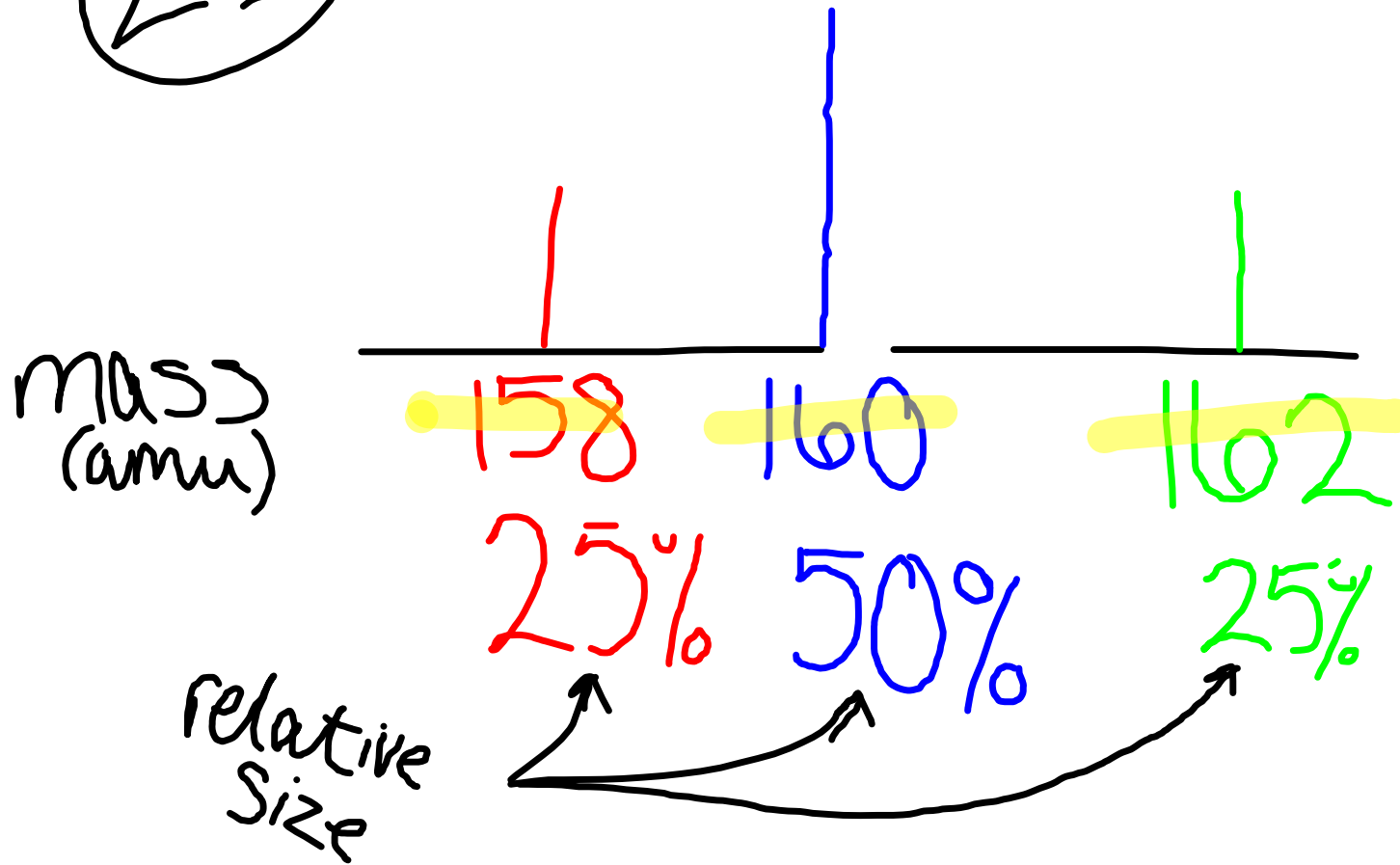
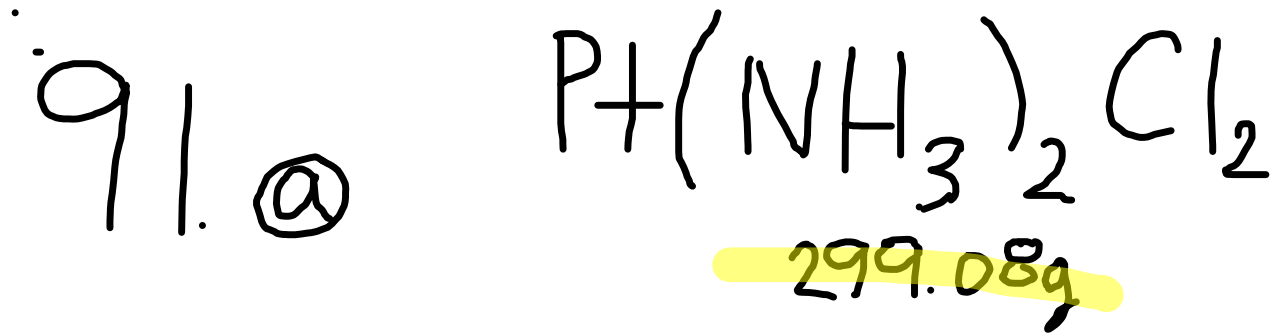


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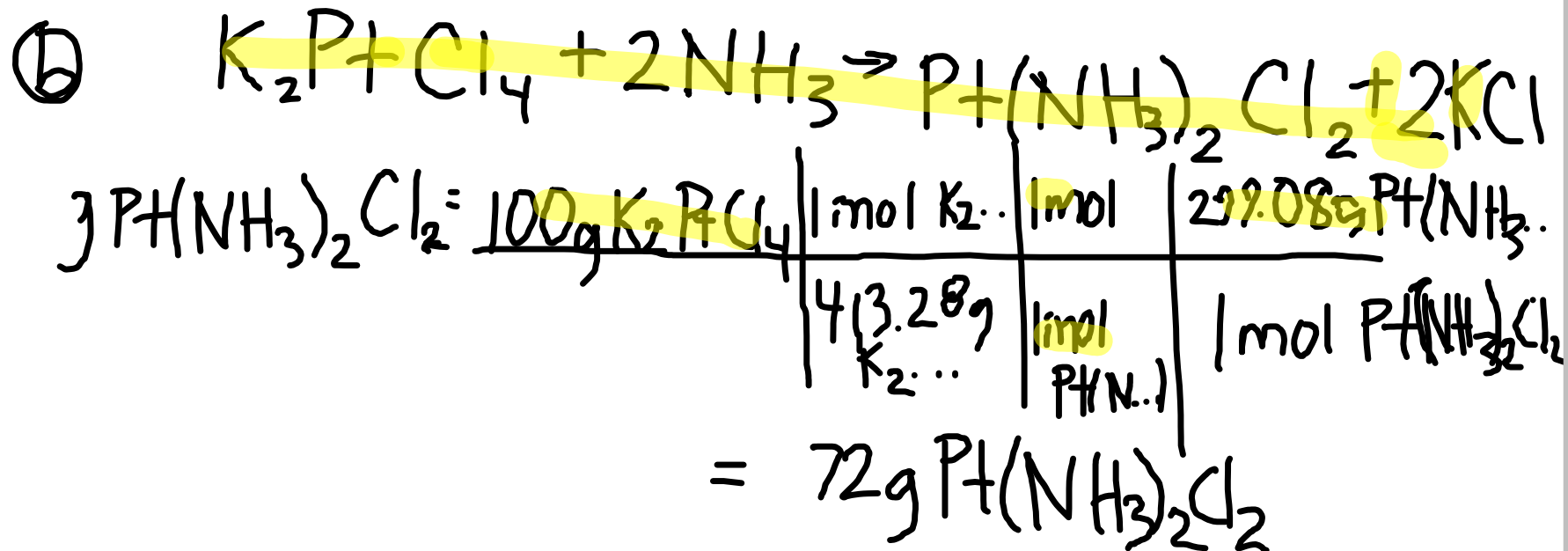
$$\text{Pt} = \frac{(195.08)}{299.08} \times 100 = 65.23\%$$

$$\text{N} = \frac{2(14)}{299.08} \times 100 \approx 2\%$$

$$\text{H} = \frac{6}{299.08} \times 100 = 23.7\%$$

$$Cl = \frac{(35.453)^2}{299.08} \times 100 = 23.7\%$$

$$\approx 100.26\%$$



$$\begin{array}{l}
 \text{g KCl} = \frac{100 \text{g K}_2\text{PtCl}_6}{1} \cdot \frac{1 \text{mol K}_2\text{Pt}_6}{413.28 \text{g K}_2\text{Pt}_6} \cdot \frac{2 \text{mol KCl}}{1 \text{mol K}_2\text{PtCl}_6} \cdot \frac{74.6 \text{g KCl}}{1 \text{mol KCl}} \\
 \\
 = 36.1 \text{g KCl}
 \end{array}$$